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strong bases Acids and

Bases, pH and pOH

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Titration Curves, pH
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15.1 Acids and bases

acid derived from Latin
acidus (meaning sour or

tart) related to Latin

acetum (meaning

vinegar) characteristic

properties associated

with acid: 1. sour taste

2. change the color of

litmus from blue to red

3. react with ?metal

(such as Zn, Mg) to

produce H₂ gas

?hydroxide base to

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produce H_2O and salt
?carbonate to produce
 CO_2

~~Chapter 15 Acids, Bases and Salts~~

Acids and Bases.

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Acids and Bases. What
we will learn: •Bronsted
acids and bases •Acid-
base properties of water
•pH - a measure of
acidity •Strength of

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Acids and Bases • Weak acids (bases) and acid (base) ionization constant • Diprotic and polyprotic acids • Molecular structure and strength of acids • Acid-base properties of salts • Acid-base properties of oxides and hydroxides • Lewis acids and bases.

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~~Bases~~ Acids Bases

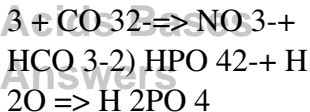
15 Drill: Identify the B-L acid and base in each of the following. Circle any amphoteric species
acid acid base base

Note: In both examples, water behaved as an acid or a base. A species that can act as an acid or a base is called

amphoteric. Bronsted-Lowry (B-L) Theory –

Cont. acid base 1) HNO

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Sec 15.3: Definitions of
Acids and Bases:

Arrhenius Definition-

Acids: a substance that
produces H^+ ions in an
aqueous solution. Bases:
a substance that

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produces OH^- ions in
aqueous solution.

*Note: under this
definition, the acid and
base combine to form
water, neutralizing each
other in the process

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Multiple Choice

Questions 1) _____ is found in carbonated beverages due to the reaction of carbon dioxide with water.

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ACIDS AND BASES

453 SECTION 15-1 O

BJECTIVES List five

general properties of

aqueous acids and

bases. Name common

binary acids and

oxyacids, given their

chemical formulas. List

five acids commonly

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used in industry and the laboratory, and give two properties of each.

Define acid and base according to

Arrhenius's theory of ionization. Explain the differences

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Chapter 15: Acids and
Bases Acids and Bases

Arrhenius Definitions:

acids - compounds that
produce an increase in
[H⁺] when dissolved in
water bases -

compounds that produce
an increase in [OH⁻]

when dissolved in water

Lewis Definitions: acids

- electron pair acceptors

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bases - electron pair
donors Brønsted-Lowry
Definitions:

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Chapter 15 main idea.

acids and bases change the color of compound indicators. ... -Lewis acid-base reaction is the formation of one or more covalent bonds between an electron-pair donor and an electron pair acceptor, although the 3 acid-base reactions differ, many compounds can be categorized as

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Chapter 15 Acids and
Bases Multiple Choice
Questions 1) _____ is
found in carbonated
beverages due to the
reaction of carbon
dioxide with water. A)
CH₃COOH; B) H₂CO₃;
C) HCOOH; D)

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C_6H_5COOH ; E)

CH_3CH_2COOH ;

Answer: B. Diff: 1 Page

Ref: 15.2 2) Which of
the following species is
amphoteric? A)

CO_3^{2-} -B) HF; C) NH_4^+ ?

D) HPO_4^{2-}

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In the chapter on acids and bases, we saw two more definitions of acids: a compound that donates a proton (a hydrogen ion, H^+) to another compound is called a Brønsted-

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Lowry acid, and a Lewis acid is any species that can accept a pair of electrons.

~~15.2 Lewis Acids and Bases—Chemistry~~
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