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Answer all the questions in complete
sentences (just restate the original
question). Show the math work for #
27-30. 1. The law of conservation of

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mass follows from the concept that _____. 2. If 4 g of element A combine with 10 g of element B, then 12 g of element A combine with _____ g of element B. 3.

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Key Concepts: Terms in this set (16)
Count the sig figs in this length.
0.05730 meters. 4. how many sig figs
are in this measurement? 8.750×10^{-2} gram. 4. round the

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measurement to 3 sig figs. write in
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Multiple Choice 1. The density of an object is

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3 \square ___ KCl + ___ O₂ a. 3 b. 5 c. 6 d. 7
e. 8 6. Calculate the mass of hydrogen
formed when 27 g of aluminum reacts
with excess hydrochloric acid
according to the balanced equation
below. $2\text{Al} + 6\text{HCl} \square 2\text{AlCl}_3 + 3\text{H}_2$ a.

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1.5 g b. 2.0 g c. 3.0 g d. 6.0 g e. 12 g

7. How many grams of nitric acid, HNO_3 , can be prepared from the reaction of 138 ...

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a) Identify and describe the proton,

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neutron and electron in terms of their relative charges and relative masses.

b) Discuss the behaviour of beams of protons, neutrons and electrons in electric fields. c) Calculate the distribution of mass and charges within an atom from the given data. d) Deduce the number of protons,

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neutrons and electrons present in both atoms and ions for a given proton and nucleon numbers/charge.

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Chapter 3 Test Multiple Choice Identify the choice that best completes the statement or answers the question. 1. Most of the volume of an atom is occupied by the a. protons. c. nucleus. b. electron cloud. d. nuclides. 2. When the light from excited atoms of an element is passed through a prism, the

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distinct colored lines produced are called 3.

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1, Practice #1; Exam 1, Practice #2;
Exam 1, Practice #3; Exam 1, Practice
#4. Exam 2 Exam 2, Practice #1;
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3 Name _____ MULTIPLE CHOICE.
Choose the one alternative that best completes the statement or answers the question. 1) Silver has two naturally occurring isotopes with the following isotopic masses: 107.471825 Ar 106.90509 Ar 108.9047 Ar The average atomic mass of silver is 107.8682

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